#

**WORKSHEET**

**GRADE: XII NAME: DATE: 27.09.19**

# SUBJECT: CHEMISTRY TOPIC: ELECTROCHEMISTRY

1. The conductivity of 0.20 M solution of KCl at 298 K is 0.0248 S cm-1. The molar conductivity is

(A) 120$Ω^{-1}$cm2 mol-1

(B) 124$Ω^{-1}$ cm2 mol-1

(C) 200$Ω^{-1}$ cm2 mol-1

(D) 75$ Ω^{-1}$cm2 mol-1

2. Reaction that takes place at graphite anode in dry cell is

(A) Zn2++2e-→Zn

(B)Zn→Zn2++2e-

(C)Mn2++2e- →Mn

(D)Mn → Mn++ e- +1.5 V.

3. The hydrogen electrode is dipped in a solution of p =3 at 25. The potential of the cell would be (the value of 2.303 RT/F is 0.059 V)

(A) 0.177 V

(B) 0.087 V

(C) -0.177 V

(D) 0.059 V

4. The standard emf of the cell, Zn+Cu2+→Cu+Zn2+ is 1.10 V at 25$℃$. The emf of the cell when 0.1 M Cu2+ and 0.1 M Zn2+ solutions are used will be

(A) 1.07 V

(B) 0.110 V

(C) -1.10 V

(D) -0.110 V

5. If a current of 0.5 ampere flows through a metallic wire for 2 hours, then how many electrons would flow through the wire?

(A) 2.25 x 10-22 electrons

(B) 3.7 x 10-11 electrons

(C) 7.5 x 1024 electrons

(D) 2.25 x 1022 electrons